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reactions can be identified by the reaction a carbon containing compound with oxygen gas to always produce water and carbon dioxide.

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Chemical Equations Flashcards ...

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Unit 7 Chemical Equations and Reactions What is a Chemical Equation? A Chemical Equation is a written representation of the process that occurs in a chemical reaction. A chemical equation is written with the Reactants on the left side of an arrow and the

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Products of the chemical reaction on the right side of the equation.

Unit 7 Chemical Equations and Reactions

Unit 7 Chemical Reactions. The solid product produced during a chemical reaction in a solution is

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called a precipitate. Water is usually formed with the combination of an acid and a base. A metal salt is also formed. Gas formed when a gas is not a reactant.

Unit 7 Chemical Reactions
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Balancing Chemical Reactions

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unchanged may be counted as a single unit. 4. Balance the elements one at a time by using coefficients. You must never attempt to balance an equation by changing the formulas. 5. Go through the equation a second time to be sure that each element and polyatomic

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ion is balanced. 6. Make sure that all coefficients are in the lowest possible ratio.

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Conservation of Mass PPT (7-8)
Complete Lab-Law of Conservation
of Mass Vodcast 7.1 (Intro to
Chemical Reactions) With notes
TUESDAY 1.6.15 OBJECTIVES IN-
CLASS AGENDA HOMEWORK
WEEKLY VIDEOS We will practice
writing the word form of a chemical

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equation. Practice Writing
Equations

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Unit 7 Balancing Equations

Worksheet 2 Answer Key ...

Unit 7: Chemical Equations Page 8

Types of Chemical Reactions

Directions (a) Write and balance the given equation. (b) Indicate the type

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- of chemical reaction represented. 1.
Iron reacts with oxygen gas to
produce Iron (III) oxide. (a) (b) 2.
Propane (C₃H₈) reacts with
oxygen gas to produce carbon
dioxide and water. (a) (b) 3.

Name: Unit 7- Chemical Equations

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Synthesis-occur when two substances (generally elements) combine and form a compound.
reactant+reactant-> products (A+B ->AB)
Decomposition-occur when a compound breaks up into the elements or in a few simpler compounds (1reactant->product+

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product) ($AB \rightarrow A+B$) Single-replacement- occur when one element replaces another in an atom.

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matter involve the rearrangement of atoms and/or the transfer of electrons and are represented using balanced chemical equations.

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Unit 7 - Chemical Reactions

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Balancing Chemical Reactions

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(Balancing)

The other plan for today was to begin Unit 7 by completing the Rearranging Atoms activity. We do not have time for the nail lab, so I figured this would be a good way to introduce balancing equations. The plan was to do the Describing

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Chemical Reactions lab tomorrow, since today's activity seemed very straight forward.

Unit 7 Chemical Reactions -
Rearranging Atoms
Chemistry: Unit 7- Chemical
Reactions. The same number and

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type of atom must be present on each side of the equation.

Balancing is accomplished by adding coefficients; never change the subscripts. Coefficients must be in the smallest whole number ratio.

Chemistry: Unit 7- Chemical

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Reactions Flashcards | Quizlet

When a solution of hydrogen chloride is added to solid sodium bicarbonate (NaHCO_3), the products are carbon dioxide, water and aqueous sodium chloride. 2.

Steam (gaseous water) reacts with carbon at high temperatures to

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produces carbon monoxide and hydrogen gases.

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Reactions- Worksheet 2
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Reactions Balancing Reaction
Equations Worksheet 1 Balance the

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following equations by inserting the proper coefficients. You must keep track of the total number of elements on each side of the equation by keeping a tally beneath each reaction. **BOX YOUR COEFFICIENTS** when finished.

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Date Pd Chemistry Unit 7 Chemical
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Unit 7 PowerPoint Slides; Language
of Chemistry Lab; Chemical vs.
Physical Reactions; Naming and
Formula Writing; Identifying 5
Types of Rxn + Activity; Mixed
Naming & Formula Writing + Sweet

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Exploration + Practice WS) Writing
& Predicting Products of a
Reaction; Unit 7 Study Guide (key)

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Balancing Chemical Reactions

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Unit 7 Overview In this unit, you will build on your model of matter as bonded atoms that combine in definite ratios to include the rearrangement of these atoms to form new substances during chemical reactions. You will continue to relate your

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observations of the behavior of matter to the microscopic structure and behavior of atoms.

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Equations Worksheet 1 Balance the following equations by inserting the proper coefficients. loading...

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Chemistry - Unit 7 Reaction

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